Topic 6: Types of Chemical Reactions



Energy:

Metabolism:

Kinetic:

Potential:

First Law of Thermodynamics:

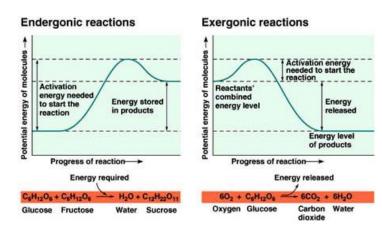
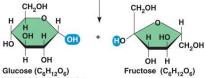


Table 1 Average Bond Energies

Bond type	Average bond energy (kJ/mol)
H-H	436
C-H	411
0-H	459
N-H	391
C-C	346
C-0	359
C=0	799
0=0	494

FORMATION OF A TIDE ROND DE 0 dehydration Suniticsis lus one molecule CH₀ CH2OH CH2OH 0 н н н но OH CH2OH НĊ он он Ĥ. Ĥ Sucrose (C12H22O11) Sucrase CH₂OH CH₂OH 0 н



Bond Energies:

Energy & Condensation/Anabolic Reactions:

Bonds Broken:

Bonds Formed:

NET

Energy & Hydrolysis/Catabolic Reactions:

Bonds Broken:

Bonds Formed:

Oxidation – Reduction (Redox) Reactions:

Coupled Redox Reactions:





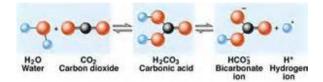
NADH



Bripad from RADH and RADH is the transformer Betrone corriers

Neutralization / Acid-Base Reactions:

Buffers:



Blood pH Levels

