

**Binary Ionic Compounds  
Divalent Cation**

Name each of the following compounds:

Formula	Name	Formula	Name
1. CuS	<b>copper(II) sulphide</b>	11. CuCl <sub>2</sub>	<b>copper(II) chloride</b>
2. PbBr <sub>4</sub>	<b>lead(IV) bromide</b>	12. CuBr	<b>copper(I) bromide</b>
3. Pb <sub>3</sub> N <sub>2</sub>	<b>lead(II) nitride</b>	13. PbO	<b>lead(II) oxide</b>
4. Fe <sub>2</sub> O <sub>3</sub>	<b>iron(III) oxide</b>	14. Fe <sub>2</sub> S <sub>3</sub>	<b>iron(III) sulphide</b>
5. FeI <sub>2</sub>	<b>iron(II) iodide</b>	15. PbCl <sub>2</sub>	<b>lead(II) chloride</b>
6. Sn <sub>3</sub> P <sub>4</sub>	<b>tin(IV) phosphide</b>	16. SnO	<b>tin(II) oxide</b>
7. Cu <sub>2</sub> S	<b>copper(I) sulphide</b>	17. Cu <sub>2</sub> O	<b>copper(I) oxide</b>
8. SnCl <sub>2</sub>	<b>tin(II) chloride</b>	18. PbO <sub>2</sub>	<b>lead(IV) oxide</b>
9. HgO	<b>mercury(II) oxide</b>	19. FeO	<b>iron(II) oxide</b>
10. Hg <sub>2</sub> F <sub>2</sub>	<b>mercury(I) fluoride</b>	20. SnO <sub>2</sub>	<b>tin(IV) oxide</b>

Write the formula for each of the following compounds:

Name	Formula	Name	Formula
1. mercury(I) oxide	<b>Hg<sub>2</sub>O</b>	11. lead(II) oxide	<b>PbO</b>
2. mercury(I) iodide	<b>HgI</b>	12. iron(III) bromide	<b>FeBr<sub>3</sub></b>
3. gold(III) chloride	<b>AuCl<sub>3</sub></b>	13. copper(II) oxide	<b>CuO</b>
4. manganese(II) oxide	<b>MnO</b>	14. mercury(I) selenide	<b>Hg<sub>2</sub>Se</b>
5. chromium(III) chloride	<b>CrCl<sub>3</sub></b>	15. lead(IV) chloride	<b>PbCl<sub>4</sub></b>
6. cobalt(II) oxide	<b>CoO</b>	16. copper(II) nitride	<b>Cu<sub>3</sub>N<sub>2</sub></b>
7. manganese (III) oxide	<b>Mn<sub>2</sub>O<sub>3</sub></b>	17. tin(IV) chloride	<b>SnCl<sub>4</sub></b>
8. cobalt(III) sulphide	<b>Co<sub>2</sub>S<sub>3</sub></b>	18. tin(IV) oxide	<b>SnO<sub>2</sub></b>
9. gold(I) fluoride	<b>AuF</b>	19. iron(III) sulphide	<b>Fe<sub>2</sub>S<sub>3</sub></b>
10. chromium(II) bromide	<b>CrBr<sub>2</sub></b>	20. iron(II) sulphide	<b>FeS</b>