Blackline Master 5.8

Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

(a) magnesium oxide	MgO	(k) copper(I) bromide	CuBr
(b) sodium fluoride	NaF	(l) tin(II) iodide	SnI_2
(c) aluminum nitride	AlN	(m) iron(III) chloride	$FeCl_3$
(d) potassium sulfide	K_2S	(n) calcium phosphide	Ca_3P_2
(e) lithium iodide	LiI	(o) lead(II) oxide	PbO
(f) calcium bromide	$CaBr_2$	(p) lead(IV) fluoride	PbF_4
(g) beryllium oxide	BeO	(q) tin(IV) bromide	$SnBr_4$
(h) nickel(II) chloride	$NiCl_2$	(r) copper(II) sulfide	CuS
(i) magnesium nitride	Mg_3N_2	(s) iron(II) oxide	FeO
(j) aluminum sulfide	Al_2S_3	(t) calcium nitride	Ca_3N_2

2. Write the names for the following compounds.

(a) Li_2O	lithium oxide	(k) PbS	lead(II) sulfide
(b) $AlCl_3$	aluminum chloride	(l) SnO_2	tin (IV) oxide
(c) MgS	magnesium sulfide	(m) Na_2S	sodium sulfide
(d) CaO	calcium oxide	(n) Mg_3P_2	magnesium phosphide
(e) KBr	potassium bromide	(o) NiO	nickel (II) oxide
(f) BeF	beryllium fluoride	(p) CuI	copper (I) iodide
(g) Na_3N	sodium nitride	(q) $PbCl_4$	lead (IV) chloride
(h) Al_2O_3	aluminum oxide	(r) FeP	iron (III) phosphide
(i) $CuCl_2$	copper (II) chloride	(s) CaF_2	calcium fluoride
(j) $FeBr_3$	iron (III) bromide	(t) K_3P	potassium phosphide