

Blackline Master 5.8

*Ionic Compounds:  
Names and Formulas Worksheet*

1. Write the formulas for the following compounds.

- |   |                              |
|---|------------------------------|
| (a) magnesium oxide _____                 | (k) copper(I) bromide _____  |
| (b) sodium fluoride _____                 | (l) tin(II) iodide _____     |
| (c) aluminum nitride _____                | (m) iron(III) chloride _____ |
| (d) potassium sulfide _____               | (n) calcium phosphide _____  |
| (e) lithium iodide _____                  | (o) lead(II) oxide _____     |
| (f) calcium bromide _____                 | (p) lead(IV) fluoride _____  |
| (g) beryllium oxide _____                 | (q) tin(IV) bromide _____    |
| (h) nickel <sup>(II)</sup> chloride _____ | (r) copper(II) sulfide _____ |
| (i) magnesium nitride _____               | (s) iron(II) oxide _____     |
| (j) aluminum sulfide _____                | (t) calcium nitride _____    |

2. Write the names for the following compounds.

- |                                   |                                   |
|-----------------------------------|-----------------------------------|
| (a) $\text{Li}_2\text{O}$ _____   | (k) $\text{PbS}$ _____            |
| (b) $\text{AlCl}_3$ _____         | (l) $\text{SnO}_2$ _____          |
| (c) $\text{MgS}$ _____            | (m) $\text{Na}_2\text{S}$ _____   |
| (d) $\text{CaO}$ _____            | (n) $\text{Mg}_3\text{P}_2$ _____ |
| (e) $\text{KBr}$ _____            | (o) $\text{NiO}$ _____            |
| (f) $\text{BeF}$ _____            | (p) $\text{CuI}$ _____            |
| (g) $\text{Na}_3\text{N}$ _____   | (q) $\text{PbCl}_4$ _____         |
| (h) $\text{Al}_2\text{O}_3$ _____ | (r) $\text{FeP}$ _____            |
| (i) $\text{CuCl}_2$ _____         | (s) $\text{CaF}_2$ _____          |
| (j) $\text{FeBr}_3$ _____         | (t) $\text{K}_3\text{P}$ _____    |