

### Ionic Compounds Practice

WRITE FORMULAS FOR EACH OF:

1. lithium nitrate \_\_\_\_\_
2. sodium chlorate \_\_\_\_\_
3. calcium nitrate \_\_\_\_\_
4. sodium phosphate \_\_\_\_\_
5. aluminum nitrate \_\_\_\_\_
6. potassium sulfate \_\_\_\_\_
7. ammonium nitrate \_\_\_\_\_
8. ammonium oxide \_\_\_\_\_
9. potassium chlorate \_\_\_\_\_
10. barium carbonate \_\_\_\_\_
11. iron (II) sulfate \_\_\_\_\_
12. iron (II) nitrate \_\_\_\_\_
13. iron (II) phosphate \_\_\_\_\_
14. titanium (IV) sulfate \_\_\_\_\_
15. nickel (III) nitrate \_\_\_\_\_
16. silver nitrate \_\_\_\_\_
17. aluminum sulfate \_\_\_\_\_
18. aluminum phosphate \_\_\_\_\_
19. copper (I) sulfate \_\_\_\_\_
20. copper (I) phosphate \_\_\_\_\_
21. zinc sulfate \_\_\_\_\_
22. cobalt (III) chloride \_\_\_\_\_
23. calcium hydroxide \_\_\_\_\_
24. sodium hydroxide \_\_\_\_\_
25. aluminum hydroxide \_\_\_\_\_
26. ammonium hydroxide \_\_\_\_\_
27. sodium bicarbonate \_\_\_\_\_
28. ammonium bicarbonate \_\_\_\_\_
29. copper (II) bicarbonate \_\_\_\_\_
30. aluminum bicarbonate \_\_\_\_\_
31. iron (III) hydrogen carbonate \_\_\_\_\_
32. sulfate radical \_\_\_\_\_

NAME EACH OF THE FOLLOWING:

33.  $\text{CaSO}_4$  \_\_\_\_\_
34.  $\text{LiOH}$  \_\_\_\_\_
35.  $\text{Ba(OH)}_2$  \_\_\_\_\_
36.  $\text{KNO}_3$  \_\_\_\_\_
37.  $\text{Pb(OH)}_2$  \_\_\_\_\_
38.  $\text{Ba}_3(\text{PO}_4)_2$  \_\_\_\_\_
39.  $(\text{NH}_4)_2\text{SO}_4$  \_\_\_\_\_
40.  $\text{Mg(NO}_3)_2$  \_\_\_\_\_
41.  $\text{Ni}_2(\text{SO}_4)_3$  \_\_\_\_\_
42.  $\text{Pb(ClO}_3)_2$  \_\_\_\_\_
43.  $\text{Na}_2\text{SO}_4$  \_\_\_\_\_
44.  $\text{SrCO}_3$  \_\_\_\_\_