

Polyatomic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.
 - a) magnesium sulphate MgSO_4
 - b) sodium chlorate NaClO_3
 - c) aluminum nitrate $\text{Al}(\text{NO}_3)_3$
 - d) potassium hydroxide KOH
 - e) lithium phosphate Li_3PO_4
 - f) calcium carbonate CaCO_3
 - g) beryllium sulphate BeSO_4
 - h) sodium bicarbonate NaHCO_3
 - i) magnesium hydroxide $\text{Mg}(\text{OH})_2$
 - j) aluminum phosphate AlPO_4
 - k) copper(I) chlorate CuClO_3
 - l) calcium sulphate CaSO_4
 - m) lead(II) nitrate $\text{Pb}(\text{NO}_3)_2$
 - n) copper(II) hydroxide $\text{Cu}(\text{OH})_2$
 - o) iron(II) phosphate $\text{Fe}_3(\text{PO}_4)_2$
 - p) calcium chlorate $\text{Ca}(\text{ClO}_3)_2$

2. Write the names for the following compounds.

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|---------------------------------|-----------------------------|
| a) Li_2CO_3 | <u>Lithium Carbonate</u> |
| b) $\text{Al}(\text{HCO}_3)_3$ | <u>Aluminum Bicarbonate</u> |
| c) $\text{Mg}_3(\text{PO}_4)_2$ | <u>Magnesium Phosphate</u> |
| d) $\text{Ca}(\text{NO}_3)_2$ | <u>Calcium Nitrate</u> |
| e) K_2SO_4 | <u>Potassium Sulfate</u> |
| g) NaNO_3 | <u>Sodium Nitrate</u> |
| h) $\text{Al}(\text{OH})_3$ | <u>Aluminum Hydroxide</u> |
| i) CuSO_4 | <u>Copper (II) Sulfate</u> |
| j) $\text{Fe}(\text{ClO}_3)_3$ | <u>Iron (III) Chlorate</u> |
| k) $\text{Pb}_3(\text{PO}_4)_2$ | <u>Lead (II) Phosphate</u> |
| l) $\text{Sn}(\text{ClO}_3)_2$ | <u>Tin (II) Chlorate</u> |
| m) NaOH | <u>Sodium Hydroxide</u> |
| p) CuNO_3 | <u>Copper (I) Nitrate</u> |
| r) FeSO_4 | <u>Iron (II) Sulfate</u> |
| s) $\text{Ca}(\text{HCO}_3)_2$ | <u>Calcium Bicarbonate</u> |
| t) K_3PO_4 | <u>Potassium Phosphate</u> |