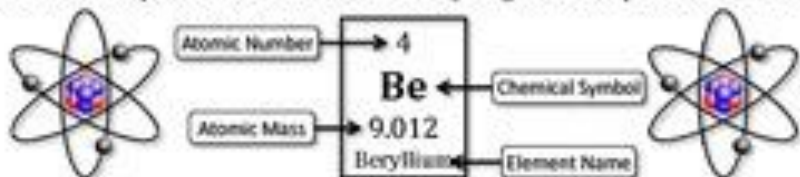


NAME _____

Periodic Table Worksheet 1

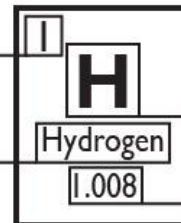
There are FOUR parts that are displayed for every element on every Periodic Table. Use this labeled picture and the Periodic Table as your guide to complete the worksheet.



DIRECTIONS: You are given ONE of the FOUR parts of the element's square – find the element using the given information and fill in the three missing parts of the square.

C	Mn	Pt	1	31	
					Vanadium
24		5	Fe		52
	Helium			Cobalt	
	Re		37	Ce	
32.065		83.798			Fermium
				87	Hf
Iridium	200.59	58.693			
			Silicon		

Atomic Number



Element Symbol



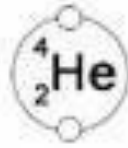
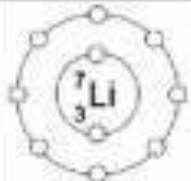

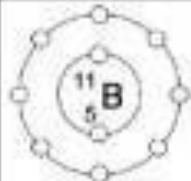
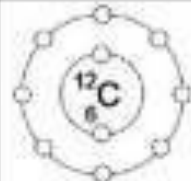
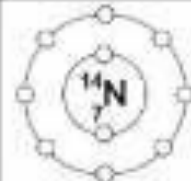
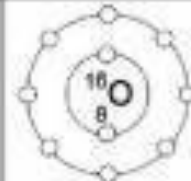
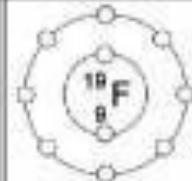
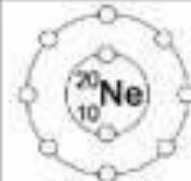
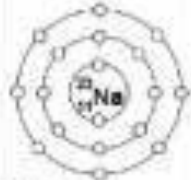


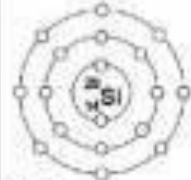
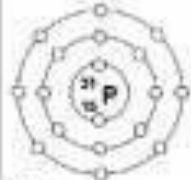
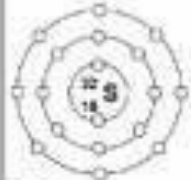
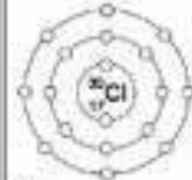
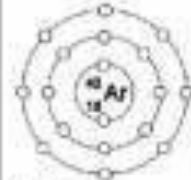
Element Name

Atomic Mass

Complete the missing information for each of the following elements from a periodic table.

Be	76		Radon		65.38	Silver
23	Nd			98.907	Am	Mercury
			Mg		186.207	46
Antimony		54.938				
		100			Nb	17
20.180	Platinum					

Why is the Periodic Table organized this way?
 Elements in the same Period have the same number of electron shells. Elements in the same Group have the same number of valence electrons - the number of electrons in their outer shell - and this gives them similar reactive properties.

	Groups 1	2	3	4	5	6	7	8
1	 Name _____ Atomic # _____ Valence e ⁻ _____	<div style="border: 1px solid black; padding: 5px;"> <p>Key</p> <p>Instructions: Give each atom the same number of electrons as protons. Color in one circle for each electron, in the numeric order given in the diagram below.</p> <p>Element Name: <u>Lithium</u></p> <p>Atomic #: <u>3</u></p> <p>Valence e⁻: <u>1</u> (Number of electrons on outer shell)</p> </div> <div style="text-align: center;">  <p>7 - Mass Number (Protons + Neutrons) Li - Chemical Symbol 3 - Atomic Number (# of Protons)</p> </div>						 Name _____ Atomic # _____ Valence e ⁻ _____
2	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____
Periods: 3	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____	 Name _____ Atomic # _____ Valence e ⁻ _____